

PII: S0957-4166(96)00442-9

Catalytic Enantioselective Conjugate Additions Using Enantiopure β -Amino Disulfides and β -Amino Thiolates

Colin L. Gibson

Department of Pure & Applied Chemistry, University of Strathclyde, 295 Cathedral Street, Glasgow, G1 1XL, UK.

Abstract: The enantioselective conjugate addition of diethylzinc to chalcone catalysed by nickel(II) using an enantiopure β -amino thiolate 2 or β -amino disulfide 1 to achieve ee's of up to 50% is described. Copyright © 1996 Elsevier Science Ltd

The asymmetric formation of carbon-carbon bonds by the use of the enantioselective conjugate addition of organometallic reagents to α,β -unsaturated carbonyl compounds is a particularly valuable synthetic method. The importance of this approach has led to considerable interest and numerous stoichiometric processes have been reported. In contrast, relatively few reports on the synthetically more useful catalytic enantioselective conjugate addition have appeared. In the area of organocopper reagents the scalemic catalysts that have been recorded thus far includes dialkylaminotropones, arenethiolates, thiosugars and amidophoshines. The enantioselective conjugate addition of dialkylamino alcohols, pyridine alcohols, amino amides, δ -amino alcohols, amino amides, amino alcohols, amino amides, amino amides, amino amides, amino amides, amino alcohols, amino amides, amino amid

Recently, we reported the beneficial effects of a sulfur chelation site on the level of asymmetric induction for bidentate amine ligands and catalysts in diorganozinc 11 and organocuprate 12 addition reactions. We anticipated that catalysts with a β -amino disulfide or a β -amino thiolate moiety would give improved enantioselectivities in the catalytic enantioselective conjugate addition of diethylzinc to chalcone 4 in the presence of nickel acetonylacetate. Herein, we report the first use of an enantiopure β -amino disulfide 1^{11} β -amino thiolate 2 in such processes to achieve ee's of up to 50%. In contrast, the corresponding β -amino alcohol, N-methylprolinol 3 gave an ee of only 6%.

Initial experiments were carried out using the thiolate 2 which was generated *in situ* by the treatment of the disulfide 1 with butyllithium in THF (see Scheme 1).¹³ The thiolate 2 was then investigated as a scalemic catalyst in the Ni(II) promoted conjugate addition of diethylzinc to chalcone 4 (see Scheme 2). In the absence of Ni(II) the conjugate addition product 5 was obtained as the racemate (entry 1, Table). Increasing the amount of Ni(II) from 1 to 7 mol% resulted in an increase in the enantioselectivity to 18 and 37% ee, respectively (entries 2 and 3). In contrast to the observations of Soai *et al.* for β -amino alcohols^{6a} no improvements in the enantioselectivity could be achieved by running the reaction in the presence of 2,2'-bipyridine (entry 4). Improved enantioselectivities were realised by the use of the disulfide 1 directly and a maximum ee of 50% was

3358 C. L. GIBSON

achieved (entry 7). The use of N-methylprolinol 3 as the scalemic catalyst gave a high yield of the addition product 5 but in only 6% ee of the (R) isomer (entry 8), this contrasts with catalysts 1 and 2 which gave the (S) enantiomer. The improvement of enantioselectivity from amino alcohol 3 (6% ee) to the disulfide 1 (50% ee) indicates the beneficial effects of a sulfur chelation site in catalysts for these processes.

Table: Enantioselective conjugate additions of diethylzinc to chalcone 4 using catalysts 1-3

Entry	Catalyst (%)	% Ni(II)	Yield (%)	% ee	config.
1	2 (20) [‡]	0	51	0	-
2	2 (17)	1	43	18	<i>(S)</i>
3	2 (17)	7	87	37	<i>(S)</i>
4	2 (17)§	7	47	34	<i>(S)</i>
5	2 (25)	7	42	16	<i>(S)</i>
6	1 (16)	3	22	44	<i>(S)</i>
7	1 (17)	7	85	50	<i>(S)</i>
8	3 (17)	7	90	6	(R)

‡ At room temperature. § With 7 mol% 2,2'-bipyridine.

REFERENCES AND NOTES

- 1. (a) Leonard, J., Contemp. Org. Synth., 1994, 1, 387; (b) Rossiter, B. E.; Swingle, N. M., Chem. Rev., 1992, 92, 771.
- 2. Ahn, K.-H.; Klassen, R. B.; Lippard, S. J., Organometallics, 1990, 9, 3178.
- 3. Lambert, F.; Knotter, D. M.; Janssen, M. D.; van Klaveren, M.; Boersma, J.; van Koten, G., Tetrahedron: Asymmetry, 1991, 2, 1097; Zhou, Q.-L.; Pfalz, A., Tetrahedron, 1994, 50, 4467.
- 4. Spescha, M.; Rihs, C., Helv. Chim. Acta, 1993, 76, 1219.
- 5. Kamai, M.; Tomioka, K., Tetrahedron Lett., 1995, 36, 4275.
- (a) Soai, K.; Hayashi, S.; Ugajin, S., Chem. Commun., 1989, 516; (b) de Vries, A. H. M.; Jansen, J. F. G. A.; Feringa, B. L., Tetrahedron, 1994, 50, 4479; (c) Fujisawa, T.; Itoh, S.; Shimizu, M., Chem. Lett., 1994, 1777.
- 7. Bolm, C.; Ewald, M.; Felder, M., Chem. Ber., 1992, 125, 1205 and references cited therein.
- 8. Corma, A.; Iglesias, M.; Martín, M. V.; Rubio, J.; Sanchez, F., Tetrahedron: Asymmetry, 1992, 3, 845
- 9. Uemura, M.; Miyake, R.; Nakayama, K.; Shiro, M.; Hayashi, Y., J. Org. Chem., 1993, 58, 1238.
- 10. Asami, M.; Usui, K.; Higuchi, S.; Inoue, S., Chem. Lett., 1994, 297.
- 11. Gibson, C. L., Chem Commun., 1996, 645
- 12. Cran, G. A.; Gibson, C. L.; Handa, S.; Kennedy, A. R., Tetrahedron: Asymmetry, 1996, 7, 2511.
- 13. Nájera, C; Yus, M; Seebach, D. Helv. Chim. Acta, 1984, 67, 289.